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CHAPTER 16 REVIEW Reaction Energy SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1. For elements in their standard state, the value of H 0 f 0 is . 2. The formation and decomposition of water can be represented by the following thermochemical equations: H 2(g) 1 2O 2(g) → H 2O(g) 241.8 kJ/mol H 2O(l) 241.8 kJ/mol → H 2(g) 1 2O 2(g)

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Title: Chapter 16 Review Reaction Energy Section 1 Short Answer Author: ï ¿ ½ ï ¿ ½Nicole Fassbinder Subject: ï ¿ ½ ï ¿ ½Chapter 16 Review Reaction Energy Section 1 Short Answer

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Chapter 16 Reaction Energy Review Answers

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REVIEW Reaction Energy

Title: Chapter 16 Reaction Energy Review Answers Author: ï ¿ ½ ï ¿ ½Thorsten Gerber Subject: ï ¿ ½ ï ¿ ½Chapter 16 Reaction Energy Review Answers

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the study of the transfers of energy as heat that accompany chemical reactions and physical changes.

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Section 16.1 Collision Theory: A Model for the Reaction Process. Goals. To describe a model, called collision theory, that helps us to visualize the process of many chemical reactions. To use collision theory to explain why not all collisions between possible reactants lead to products. To use collision theory to explain why possible reactants must collide with an energy equal to or above a certain amount to have the possibility of reacting and forming products.

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