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Chapter 9: Chemical Names and Formulas. STUDY. PLAY. Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -).

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Determined by the amount of electrons lost or gained.

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Steele. STUDY. PLAY. Monatomic ion.
- single atom with a positive or
negative charge resulting from the
loss or gain of electrons, respectively.

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Section 9.4 – Naming and Writing

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Formulas for Acids and Bases. An acid is a compound that produces H^+ ions when it dissolves in water. The formula for an acid normally starts with an H. When naming acids, you should first determine whether or not the acid contains oxygen.

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Chapter 9 Chemistry Chemical Names
and Formulas. Acids. base. Cation.
Ionic compounds. Compounds that
contain one or more hydrogen atoms
and produce.... An ionic compound

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that produces hydroxide ions when dissolved.... Any atom or group of atoms that has a positive charge. Compounds composed of cations and anions.

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Chapter 9 Chemical Names and Formulas 83 SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds.
Naming Binary Molecular Compounds

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(pages 268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

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AND FORMULAS 9
SECTION 9.2 NAMING AND WRITING
FORMULAS FOR IONIC COMPOUNDS I.~~

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Write the formulas for these binary ionic compounds. c. potassium iodide

2. write the formulas for the

compounds formed c. b. f. 3. Name

the following binary ionic compounds.

sodium sulfide g. CuCl : h. SnCl_2 a.

MnO_2 c. d. SrBr_2 221 Chapter 9

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Formulas 9.1 Naming Ions 9.2 Naming
and Writing Formulas for Ionic
Compounds 9.3 Naming and Writing
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9.4 Naming and Writing Formulas for
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Chapter 9 “ Chemical Names and
Formulas ” ... in parenthesis after the

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name of the metal (Table 9.2, p.255)

Predicting Ionic Charges Some of the
post-transition elements also have
more than one possible oxidation
state. Tin (II) = Sn^{2+} Lead (II) = Pb^{2+}
Tin (IV) = Sn^{4+} Lead (IV) = Pb^{4+}

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CHAPTER 9 Assessment 42. a. 2– b.

1+ c. 1– d. 3+ 43. a. 2+ b. 2+ c. 3+

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d. 1+ 44. a. barium ion b. iodide ion c.
silver ion d. mercury(II) ion 45.
cyanide, CN^- and hydroxide, OH^-
46. a. hydroxide ion b. lead(IV) ion c.
sulfate ion d. oxide ion 47. zero; A
compound is electrically neutral. 48.
The symbols for the cation and

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Chapter 9 Practice Problems:

Chemical Names and Formulas.

Section 9.1: Naming Ions. 1. What is the charge on the ion typically formed by each element? a. oxygen c. sodium e. nickel, two electrons lost . b. iodine d. aluminum f. magnesium.

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2. How many electrons does the neutral atom gain or lose when each ion forms? a. Cr^{3+} c. Li^+ e. Cl^- b. P^{3-}

...

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activities; ... In samples of any
chemical compound, the masses of

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the elements are always in the same proportions any atom or group of atoms that has a negative charge ...
place cation name first followed by the anion name: naming polyatomic ...

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