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Theoretical \u0026 Percent

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Stoichiometry with Mass:

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2 Molarity, Solution

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Reactants and Percent Yield

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Problem (Advanced)

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2020 Some of the worksheets below are Stoichiometry Worksheets with Answer Keys, definition of stoichiometry with tons of interesting examples and exercises involving with step by step solutions with several

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Stoichiometry: Limiting

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Limiting reagent

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Practice Problems:

Stoichiometry (Answer Key)

Balance the following

chemical reactions: a. 2CO

$+ \text{O}_2 \rightarrow 2 \text{CO}_2$. b. $2 \text{KNO}_3 \rightarrow 2$

$\text{KNO}_2 + \text{O}_2$. c. $2 \text{O}_3 \rightarrow 3 \text{O}_2$.

d. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + 2 \text{H}_2\text{O}$

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O. e. $4 \text{ CH}_3 \text{ NH}_2 + 9 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 10 \text{ H}_2 \text{ O} + 2 \text{ N}_2$. f.

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Practice Problems (Chapter 5): Stoichiometry. Practice Problems (Chapter 5): Stoichiometry. CHEM 30A. Part I: Using the conversion factors in your tool box. g A mol A mol A. 1. How many moles CH. 3. OH are in 14.8

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~~Practice Problems (Chapter
5): Stoichiometry~~

Practice Problems:

Stoichiometry. Balance the
following chemical

reactions: Hint a. $\text{CO} + \text{O}_2$

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CO₂ b. KNO₃ KNO₂ + O₂ c.
O₃ O₂ d. NH₄ NO₃ N₂ O +
H₂ O e. CH₃ NH₂ + O₂ CO
2 + H₂ O + N₂ Hint f.
Cr(OH)₃ + HClO₄ Cr(ClO₄)
3 + H₂ O; Write the
balanced chemical equations
of each reaction: a. Calcium

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carbide (CaC_2) reacts with water to form calcium hydroxide ($\text{Ca}(\text{OH})_2$) and acetylene gas (C_2H_2). b.

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~~Stoichiometry~~

Stoichiometry example

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problem 1. Stoichiometry

example problem 2. Practice:

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Practice: Converting moles
and mass. Next lesson.

Limiting reagent
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Extra Stoichiometry Problems

1. Silver nitrate reacts with barium chloride to form silver chloride and barium nitrate. a. Write and

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Problems With Answers
balance the chemical

equation. $2 \text{AgNO}_3 + \text{BaCl}_2 \rightarrow$

$2 \text{AgCl} + \text{Ba}(\text{NO}_3)_2$ b. If

39.02 grams of barium

chloride are reacted in an

excess of silver nitrate,

how many.

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Answers: Moles and

Stoichiometry Practice

Problems 1) How many moles
of sodium atoms correspond
to 1.56×10^{21} atoms of
sodium? 1.56×10^{21} atoms

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Na x 1 mol Na = 2.59 x 10³
mol Na 236.022 x 10 atoms Na

2) Determine the mass in grams of each of the following:

a. 1.35 mol of Fe
1.35 mol Fe x 55.845 g Fe =
75.4 g Fe 1 mol Fe

b. 24.5 mol O

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Problems~~

Solution Stoichiometry
Worksheet Solve the
following solutions
Stoichiometry problems: 1.

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Problems With Answers

How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate?

$$2 \text{ AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{ KNO}_3(\text{aq})$$

0.150 L AgNO_3

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0.500 moles AgNO_3 1 moles
 Ag_2CrO_4 331.74 g Ag_2CrO_4

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Worksheet~~

Stoichiometry: Mass-Mass
Problems. Show all work in

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dimensional analysis and

include correct units.

$2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$. How many grams of potassium chloride, KCl , are produced if 25.0g of potassium chlorate, KClO_3 , decompose?

$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$. How many

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grams of hydrogen, H_2 , are necessary to react completely with . 50.0 g of nitrogen, N_2 ?

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Worksheet Solve the Problems With Answers

following stoichiometry
grams-grams problems: 1)
Using the following
equation: $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$
How many
grams of sodium sulfate will
be formed if you start with

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200.0 grams of sodium

hydroxide and you have an
excess of sulfuric acid? 2)

Using the following
equation:

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Worksheet~~

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Practice Problems: Writing and classifying equations; Answers. Practice balancing chemical equations (interactive) Click "Balancing Chemical Equations Tutorial" on the left. From the Chem Team:

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Worksheet of mass mole conversions Answers to Worksheet of mass mole conversions. Here's a tutorial from ChemTutor on classifying and balancing chemical equations with Practice Problems on the

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2~~

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reading, it will be better.

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Problem : $2\text{Al} + 3\text{Cl}_2 \rightarrow 2\text{AlCl}_3$

3 When 80 grams of aluminum
is reacted with excess

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chlorine gas, how many
formula units of AlCl_3 are
produced? $\times 1 \text{ mole Al} = 2.96$
moles Al : There is a 1:1
ratio between Al and AlCl_3 ,
therefore there are 2.96
moles AlCl_3 . = 1.78×10^{25}

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Do not forget to do the stoichiometry relay questions from the previous class. Answers for those will also be posted on my website with full solutions.

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Sample questions: Multiple
Choice 1. Which one of the
following is a definition of
a chemical reaction? a)
Making new bonds between
atoms as the old ones are
broken

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